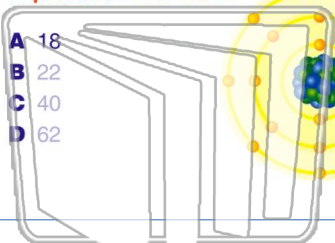




Name _____ Class _____ Date _____

1

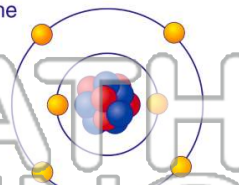
An atom contains **22 neutrons** and **40 nucleons**. What is the total number of **protons** in the atom?



2

Which **principal quantum number** is assigned to the valence electrons of a **carbon atom** in the ground state?

- A 1
B 2
C 3
D 4



3

Compared to the maximum number of electrons that can occupy the d sublevel, the **maximum number of electrons** that can

4

What was the **original mass** of a radioactive sample that decayed to **25 grams** in **four half-life periods**?

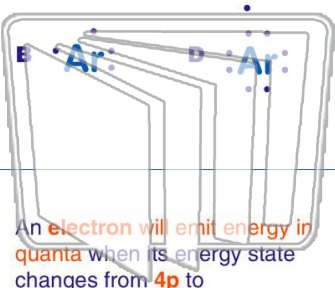
5



PREVIEW

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7



9

An **electron** will emit energy in **quanta** when its energy state changes from **4p** to

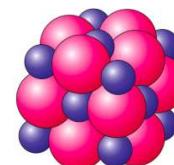
- A 5s
B 5p
C 3s
D 6p

10

- A 49 electrons, 49 protons, and 85 neutrons
B 49 electrons, 49 protons, and 49 neutrons
C 36 electrons, 36 protons, and 85 neutrons
D 36 electrons, 36 protons, and 49 neutrons

If the **electronegativity difference** between the elements in compound **NaX** is **2.0**, what is element **X**?

- A bromine
B chlorine
C fluorine
D oxygen



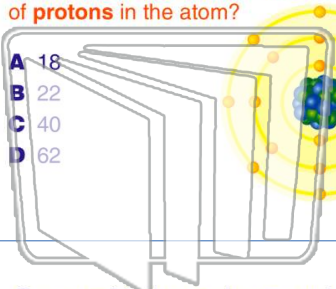


Name _____ Class _____ Date _____

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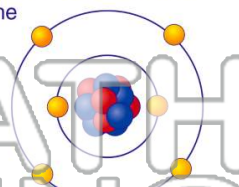
- A 18
- B 22
- C 40
- D 62



2

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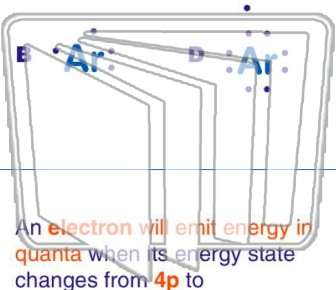
5



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