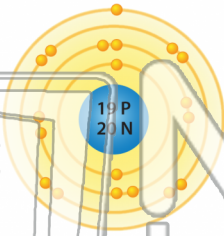




Name _____ Class _____ Date _____

1 Label the parts of the atom.

- A nucleus
- B lowest energy level
- C highest energy level
- D valence electron



2 Label the parts of the dot diagram of potassium bromide.



- A positive ion
- B negative ion
- C charge of the metal ion
- D charge of the nonmetal ion

3 Match the term with the correct description.

- ionic bond
- A. bond that forms when electrons are

4 Match the term with the correct description.

- positive ion
- A. atom with more protons than electrons

5



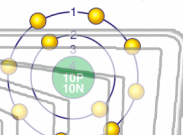
PREVIEW

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7

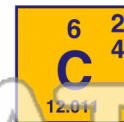
neutral. Using the diagram below, determine the **overall charge**.

- A positive
- B negative
- C neutral
- D static



does carbon need to have a **full valence shell**?

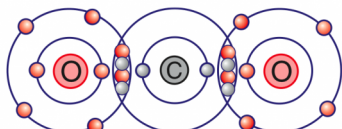
- A 2 more
- B 4 more
- C 8 more
- D 0 more, it is already full



9

Using this diagram of carbon dioxide, determine how many **valence electrons** each **oxygen** atom has after bonding with the carbon atom.

- A 8
- B 6
- C 4
- D 2



10

When a group of **active metal atoms** are put together, a strong force of **attraction** called a(n) _____ develops between these atoms.

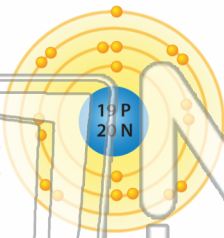
- A ionic bond
- B covalent bond
- C metallic bond
- D magnetic bond



Name _____ Class _____ Date _____

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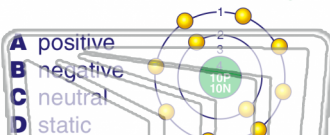


PREVIEW

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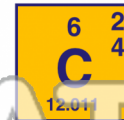
7

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does carbon need to have a **full valence shell**?

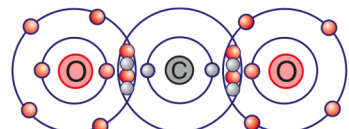
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