

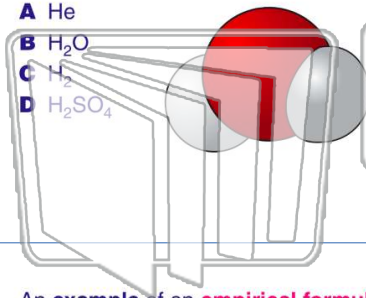


Name _____ Class _____ Date _____

1

An **example** of a **binary compound** is

- A He
- B H_2O
- C H_2
- D H_2SO_4



2

Which formula is **correct** for **ammonium sulfate**?

- A NH_4SO_4
- B $(NH_4)_2SO_4$
- C $NH_4(SO_4)_2$
- D $(NH_4)_2(SO_4)_2$



3

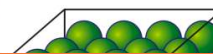
An **example** of an **empirical formula** is

- A CH_4

4

Which **compound** is a **network solid** at STP?

- A CO_2



5

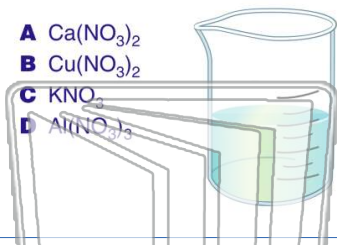


PREVIEW

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7

- A $Ca(NO_3)_2$
- B $Cu(NO_3)_2$
- C KNO_3
- D $Al(NO_3)_3$



is **formed between** these two atoms?

- A X loses electrons to Y to form an ionic bond.
- B X loses electrons to Y to form a covalent bond.
- C X gains electrons from Y to form an ionic bond.
- D X gains electrons from Y to form a covalent bond.

9

A compound consists of **25.9% nitrogen** and **74.1% oxygen** by mass. What is the **empirical formula** of the compound?

- A NO
- B NO_2
- C N_2O
- D N_2O_5

10

Which **substance** is a **binary compound**?

- A barium
- B beryllium
- C carbon dioxide
- D calcium hydroxide



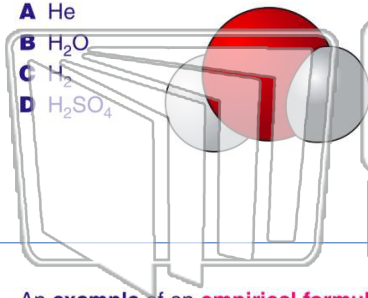


Name _____ Class _____ Date _____

1

An **example** of a **binary compound** is

- A He
- B H₂O
- C H₂
- D H₂SO₄



2

Which formula is **correct** for **ammonium sulfate**?

- A NH₄SO₄
- B (NH₄)₂SO₄
- C NH₄(SO₄)₂
- D (NH₄)₂(SO₄)₂



3

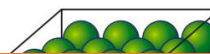
An **example** of an **empirical formula** is

- A CH₄

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Which **compound** is a **network solid** at STP?

- A CO



5

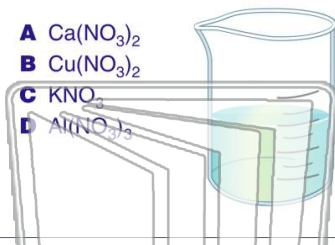


PREVIEW

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7

- A Ca(NO₃)₂
- B Cu(NO₃)₂
- C KNO₃
- D Al(NO₃)₃



is **formed between** these two atoms?

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- B NO₂
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- D N₂O₅

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