

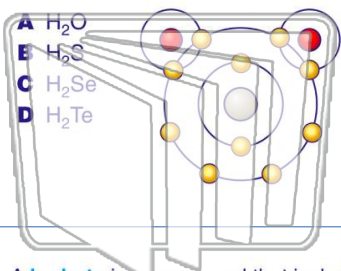


Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

1

Which of these substances has the **strongest intermolecular forces**?

- A H<sub>2</sub>O
- B H<sub>2</sub>S
- C H<sub>2</sub>Se
- D H<sub>2</sub>Te



2

Which substance contains **metallic bonds**?

- A Hg(l)
- B H<sub>2</sub>O(l)
- C NaCl(s)
- D C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>(s)

NEW PATH LEARNING

3

A **hydrate** is a compound that includes water molecules within its crystal structure. During an experiment to determine the percent by mass of water in a hydrated crystal, a student

4

In **saturated hydrocarbons**, carbon atoms are **bonded** to each other by

A single covalent bonds only

5



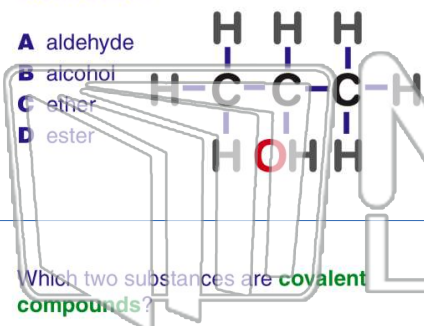
## PREVIEW

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7

shown below?

- A aldehyde
- B alcohol
- C ether
- D ester



- A FeP
- B Fe<sub>3</sub>P<sub>2</sub>
- C FePO<sub>4</sub>
- D Fe<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>

NEW PATH LEARNING

9

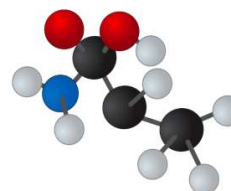
Which two substances are **covalent compounds**?

- A C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>(s) and KI(s)
- B C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>(s) and HCl(g)
- C KI(s) and NaCl(s)
- D NaCl(s) and HCl(g)

10

Which type of **change** must occur to form a **compound**?

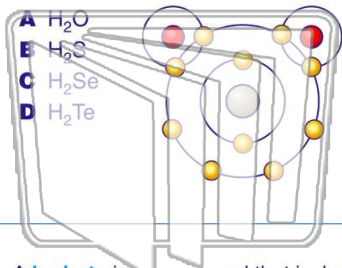
- A chemical
- B physical
- C nuclear
- D phase





Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

1 Which of these substances has the **strongest intermolecular forces**?



2 Which substance contains **metallic bonds**?

- A  $\text{Hg(l)}$   
 B  $\text{H}_2\text{O(l)}$   
 C  $\text{NaCl(s)}$   
 D  $\text{C}_6\text{H}_{12}\text{O}_6\text{(s)}$

3 A **hydrate** is a compound that includes water molecules within its crystal structure. During an experiment to determine the percent by mass of water in a hydrated crystal, a student

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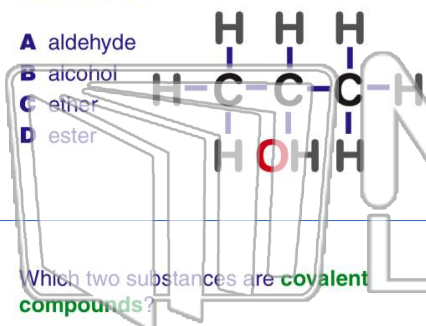
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7

shown below?

- A aldehyde  
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- A  $\text{FeP}$   
 B  $\text{Fe}_3\text{P}_2$   
 C  $\text{FePO}_4$   
 D  $\text{Fe}_3(\text{PO}_4)_2$

9 Which two substances are **covalent compounds**?

- A  $\text{C}_6\text{H}_{12}\text{O}_6\text{(s)}$  and  $\text{KI(s)}$   
 B  $\text{C}_6\text{H}_{12}\text{O}_6\text{(s)}$  and  $\text{HCl(g)}$   
 C  $\text{KI(s)}$  and  $\text{NaCl(s)}$   
 D  $\text{NaCl(s)}$  and  $\text{HCl(g)}$

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- A chemical  
 B physical  
 C nuclear  
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