

Chemical Equations



Name Class Date Which formula correctly represents the Given the equation: product of an addition reaction between $C(s) + H_2O(g) \rightarrow CO(g) + H_2(g)$ ethene and chlorine? Which species undergoes reduction? A CH GIZ C(s)CH36L В C2H₄Cl₂ C^2 D C₂H₃Cl 3 Given the unbalanced equation: Given the reaction: $Fe_2O_3 + CO \rightarrow Fe + CO_2$ 2 Al(s) + Fe₂O₃(s) $\stackrel{\text{heat}}{\longrightarrow}$ Al₂O₃(s) + 2 Fe(s) 5 **PREVIEW** Please Sign In or Sign Up to download the printable version of this worksheet 7 A T OI2 - O211501 T 1101 Which molecule is represented by X? $A H_2O(g) \to H_2O(I)$ $C(s) + O_2(g) \rightarrow C_2O(g)$ A C₂H₄ B C.H. D C₂H₈ Civen the balanced equation 9 $KNO_3(s) + 34.89 \text{ kJ} \xrightarrow{H_2O} K^+(aq) + NO_3^-(aq)$ NH₃ + HCI → NH₄CI In this reaction, ammonia molecules (NH₃) Which statement best describes this process? act as a base because they A It is endothermic and entropy increases. A accept hydrogen ions (H+) B It is endothermic and entropy decreases. **B** accept hydroxide ions (OH-) C It is exothermic and entropy increases. C donate hydrogen ions (H+) **D** It is exothermic and entropy decreases. D donate hydroxide ions (OH-)



Chemical Equations



Name_	CI	assDate	9
1	Which formula correctly represents the product of an addition reaction between ethene and chlorine? CH ₃ Cl C ₂ CH ₄ Cl ₂ C ₂ H ₃ Cl	Given the equation:	
3	Given the unbalanced equation: _Fe ₂ O ₃ + _CO → _Fe + _CO ₂ When the equation is correctly belonged	Given the reaction: $2 \text{ Al(s)} + \text{Fe}_2\text{O}_3(\text{s}) \xrightarrow{\text{heat}} \text{Al}_2\text{O}_3(\text{s})$	+ 2 Fe(s)
5	PREV Please Sign In or Sign the printable version	gn Up to download	C
7	Which molecule is represented by X? A C ₂ H C ₂ H	$\begin{array}{c} \textbf{A} \ \text{H}_2\text{O}(g) \rightarrow \text{H}_2\text{O}(I) \\ \textbf{B} \ \text{C}(s) + \text{O}_2(g) \rightarrow \text{C}_2\text{O}(g) \\ \textbf{C} \ _1^2 \text{H} + _1^3 \text{H} \rightarrow _2^4 \text{He} + _0^1 \text{n} \\ \textbf{D} \ _{92}^{235} \text{U} + _0^1 \text{n} \rightarrow _{56}^{142} \text{Ba} + _{36}^{91} \text{Kr} + \\ \end{array}$	C