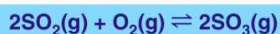




Name _____ Class _____ Date _____

1

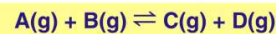
Given the reaction:

The value of the equilibrium constant (K_{eq}) will be changed by increasing the

- A pressure
- B temperature
- C concentration of $\text{SO}_2(\text{g})$
- D concentration of $\text{SO}_3(\text{g})$

2

Given the reaction at equilibrium:

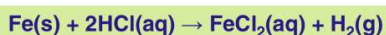


The addition of a catalyst will

- A shift the equilibrium to the right
- B shift the equilibrium to the left
- C increase the rate of forward and reverse reactions equally
- D have no effect on the forward or reverse reactions

3

Given the reaction:



In this reaction, 5 grams of powdered iron

4

Given the reaction:

If the concentration of the $\text{HCl}(\text{aq})$ is increased, the frequency of reacting collisions will

5



PREVIEW

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7

At which **temperature** will the reaction occur at the **greatest rate**?

- A 25°C
- B 50°C
- C 75°C
- D 100°C

An **increase** in the concentration of $\text{A}_2(\text{g})$ will

- A decrease the production of $\text{AB}(\text{g})$
- B decrease the frequency of collisions between $\text{A}_2(\text{g})$ and $\text{B}_2(\text{g})$
- C increase the production of $\text{B}_2(\text{g})$
- D increase the frequency of collisions between $\text{A}_2(\text{g})$ and $\text{B}_2(\text{g})$

9

Given the reaction:



In the reaction, the oxidizing agent is

- A $\text{F}_2(\text{g})$
- B $\text{Br}^{-}(\text{aq})$
- C $\text{Br}_2(\text{l})$
- D $\text{F}^{-}(\text{aq})$

10

Which **process** occurs in an operating **electrochemical cell**?

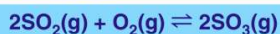
- A a reduction reaction, only
- B an oxidation reaction, only
- C a chemical reaction produced by an electric current
- D a chemical reaction that produces an electric current



Name _____ Class _____ Date _____

1

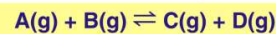
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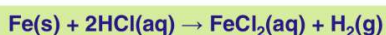


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