



Name _____ Class _____ Date _____

1 As elements of **Group 15** of the Periodic Table are considered in order from top to bottom, the **metallic character** of the atoms of each successive element generally

A decreases
B increases
C remains the same

7	N
15	P
33	As
51	Sb
83	Bi

2 Which element has a total of **5 valence electrons** present in the **fifth energy level**?

A Sb
B Bi
C I
D Br

3 For which element is the **ionic radius larger** than the **atomic radius**?

A Na

4 Which of these elements has an atom with the **most stable** outer electron configuration?

5

PREVIEW

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7 The **atomic radius** of each successive element increases. This **increase** is primarily due to an increase in

A atomic number
B mass number
C the number of protons occupying the nucleus
D the number of occupied principal energy levels

16	S
34	Se
52	Te
84	Po

8 The **metallic character** of each successive element

A decreases
B increases
C remains the same

12	Mg
20	Ca
38	Sr
56	Ba
88	Ra

9 Which elements combine by forming an **ionic bond**?

A sodium and potassium
B sodium and oxygen
C carbon and oxygen
D carbon and sulfur

10 Which element forms an ion that is **larger** than its atom?

A aluminum
B chlorine
C magnesium
D sodium



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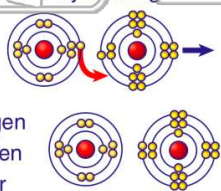
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