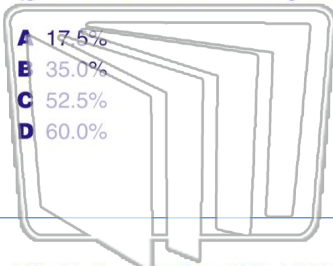Name_____ Class_____ Date_____

1 What is the **percent composition** by mass of **nitrogen** in **NH_4NO_3** (gram-formula mass = 80.0 grams/mole)?

A 17.5%
B 35.0%
C 52.5%
D 60.0%

2 How many milliliters of 0.100 M NaOH(aq) would be needed to **completely neutralize** 50.0 milliliters of 0.300 M HCl(aq)?

A 16.7 mL
B 50.0 mL
C 150 mL
D 300 mL

3 What is the **density** of **N_2** at STP?

A 1.00 g/L

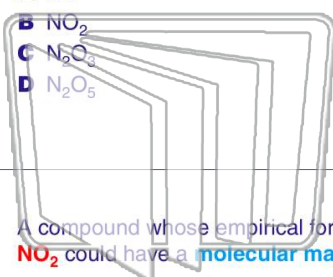**4** Which concentration of a sugar solution has a **boiling point of 100.52°C** at standard pressure?

5

PREVIEW

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7 and **69.6% oxygen** by mass?
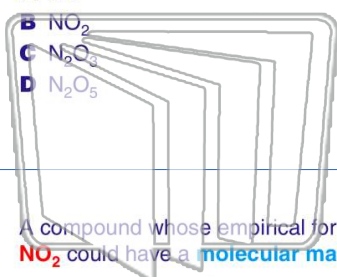
A NO
B NO_2
C N_2O_3
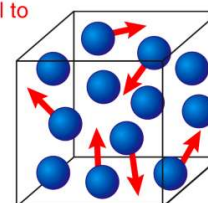D N_2O_5

A 15%
B 20%
C 35%
D 60%

7
N
Nitrogen
14.007

9 A compound whose empirical formula is **NO_2** could have a **molecular mass** of

A 23
B 39
C 92
D 120

10 The density of a gas is **1.43 grams per liter** at STP. The mass of **1 mole** of this gas is equal to
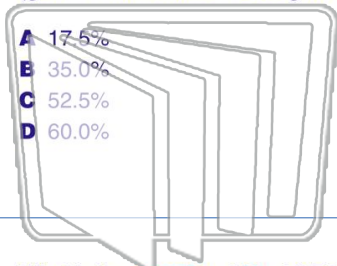
A 1.43 g
B 15.7 g
C 22.4 g
D 32.0 g

Name_____ Class_____ Date_____

1 What is the **percent composition** by mass of **nitrogen** in **NH₄NO₃** (gram-formula mass = 80.0 grams/mole)?

A 17.5%
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C 52.5%
D 60.0%

B

2 How many milliliters of 0.100 M NaOH(aq) would be needed to **completely neutralize** 50.0 milliliters of 0.300 M HCl(aq)?

A 16.7 mL
B 50.0 mL
C 150 mL
D 300 mL

C

3 What is the **density** of N_2 at STP?

A 1.00 g/L

**4** Which concentration of a sugar solution has a **boiling point of 100.52°C** at standard pressure?

A

5

D

7 and 69.6% oxygen by mass?

A NO
B NO_2
C N_2O_3
D N_2O_5

B

A 15%
B 20%
C 35%
D 60%

7
N
Nitrogen
14.007

C

9 A compound whose empirical formula is **NO₂** could have a **molecular mass** of

A 23
B 39
C 92
D 120

C

10 The density of a gas is **1.43 grams per liter** at STP. The mass of **1 mole** of this gas is equal to

A 1.43 g
B 15.7 g
C 22.4 g
D 32.0 g

D