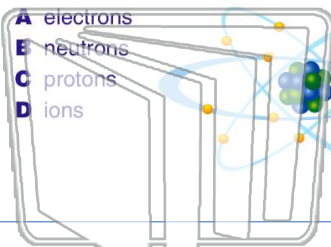




Name _____ Class _____ Date _____

1 When a **redox reaction** occurs, there must be a **transfer** of

- A electrons
- B neutrons
- C protons
- D ions



2 What is the **oxidation number** of **carbon** in NaHCO_3 ?

- A -2
- B +2
- C -4
- D +4



3 A **redox reaction** is set up so that both half reactions take place in separate beakers that are connected by a salt bridge and an external conductor. A **path for the transfer**

4 An **oxidation half-reaction** always **involves** the
A gain of electrons and a decrease in

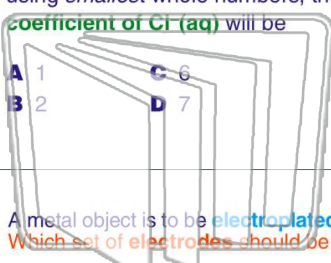


PREVIEW

Please [Sign In](#) or [Sign Up](#) to download the printable version of this worksheet

7 When the equation is correctly balanced using *smallest* whole numbers, the **coefficient of $\text{Cl}^- (\text{aq})$** will be

- A 1
- B 2
- C 6
- D 7



10 What is the **net cell potential (E^\ominus)** for the overall reaction?

- A 0.45 V
- B 1.92 V
- C 2.37 V
- D 2.82 V

9 A metal object is to be **electroplated** with silver. Which **set of electrodes** should be used?

- A a silver anode and a metal object as the cathode
- B a platinum anode and a metal object as the cathode
- C a silver cathode and a metal object as the anode
- D a platinum cathode and a metal object as the anode



10 In an **electrolytic cell**, the **anode** is **always** the

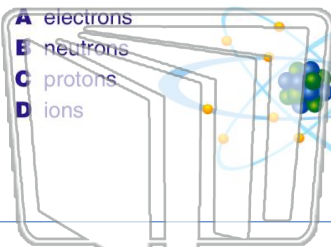
- A negative electrode, where reduction occurs
- B negative electrode, where oxidation occurs
- C positive electrode, where reduction occurs
- D positive electrode, where oxidation occurs



Name _____ Class _____ Date _____

1 When a **redox reaction** occurs, there must be a **transfer** of

- A electrons
- B neutrons
- C protons
- D ions



2 What is the **oxidation number** of **carbon** in NaHCO_3 ?

- A -2
- B +2
- C -4
- D +4



3 A **redox reaction** is set up so that both half reactions take place in separate beakers that are connected by a salt bridge and an external conductor. A **path for the transfer**

4 An **oxidation half-reaction** always **involves** the

A gain of electrons and a decrease in

5



PREVIEW

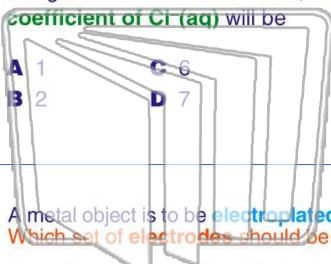
Please [Sign In](#) or [Sign Up](#) to download the printable version of this worksheet

7



When the equation is correctly balanced using *smallest* whole numbers, the **coefficient of $\text{Cl}(\text{aq})$** will be

- A 1
- B 2
- C 6
- D 7



What is the **net cell potential (E^\ominus)** for the overall reaction?

- A 0.45 V
- B 1.92 V
- C 2.37 V
- D 2.82 V

9 A metal object is to be **electroplated** with silver. Which **set of electrodes** should be used?

- A a silver anode and a metal object as the cathode
- B a platinum anode and a metal object as the cathode
- C a silver cathode and a metal object as the anode
- D a platinum cathode and a metal object as the anode



10 In an **electrolytic cell**, the **anode** is always the

- A negative electrode, where reduction occurs
- B negative electrode, where oxidation occurs
- C positive electrode, where reduction occurs
- D positive electrode, where oxidation occurs