Oxidation, Reduction and Electrochemistry

CHEM

Name_____ Class_____ Date_____

1 In an **oxidation-reduction reaction**, **reduction** is defined as the

A loss of protons
B gain of protons
C loss of electrons
D gain of electrons

2 What is the **oxidation number** assigned to **manganese** in $KMnO_4$?

A +7
B +2
C +3
D +4

3 Which of the following **aqueous solutions** is the best **conductor of electricity**?

4 A **chemical cell** is made up of two half-cells connected by an external conductor and a salt bridge. The **function** of the **salt bridge** is to

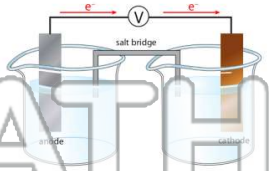
5

7

A $Fe(s) \rightarrow Fe^{2+}(aq) + 2e^-$
B $Fe^{2+}(aq) \rightarrow Fe(s) + 2e^-$
C $Fe(s) + 2e^- \rightarrow Fe^{2+}(aq)$
D $Fe^{2+}(aq) + 2e^- \rightarrow Fe(s)$

supplied by the cell

A decreases
B increases
C remains the same

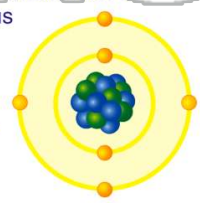


9 In an **oxidation-reduction reaction**, the **oxidation number** of the **oxidizing agent**

A decreases
B increases
C remains the same

10 As ^{14}C decays to ^{14}N, the number of **protons** in the nucleus

A decreases
B increases
C remains the same



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1 In an **oxidation-reduction reaction**, **reduction** is defined as the

A loss of protons
B gain of protons
C loss of electrons
D gain of electrons

D

2 What is the **oxidation number** assigned to **manganese** in $KMnO_4$?

A +7
B +2
C +3
D +4

A

3 Which of the following **aqueous solutions** is the best **conductor of electricity**?

A 0.10 M CH_3OH

4 A **chemical cell** is made up of two half-cells connected by an external conductor and a salt bridge. The **function** of the **salt bridge** is to

A

5

D

7

A $Fe(s) \rightarrow Fe^{2+}(aq) + 2e^-$
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supplied by the cell

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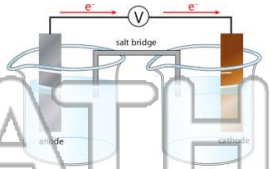

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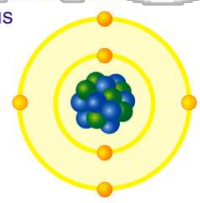

B