

## Oxidation, Reduction and Electrochemistry



Name_	Cla	ass	Date	
1	In an <b>oxidation-reduction reaction</b> , reduction is defined as the		ne oxidation number assigned nese in KMnO <sub>4</sub> ?	
	A loss of protons  C loss of electrons  D gain of electrons	B + 2   1   1   1   1   1   1   1   1   1		
3	Which of the following aqueous solutions is the best conductor of electricity?	connected	I cell is made up of two half-cells by an external conductor and a salt e function of the salt bridge is to	
	A D 10 M CH OH			
		ABC		
5				
	PREVIEW			
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	the printable version of this worksheet			
7				
	<b>A</b> Fe(s) $\rightarrow$ Fe <sup>2+</sup> (aq) + 2e <sup>-</sup> <b>B</b> Fe <sup>2+</sup> (aq) $\rightarrow$ Fe(s) + 2e <sup>-</sup>	supplied I	ises	
	C Fe(s) + 2e <sup>-</sup> → Fe <sup>2+</sup> (aq) D Fe <sup>2+</sup> (aq) + 2e Fe(s)	B increa c re mair same		
9	n an oxidation-reduction reaction, the oxidation number of the oxidizing agent		cays to <sup>14</sup> N, the number of n the nucleus	
	A decreases B increases C remains the same	A decrea B increas C remain		



## Oxidation, Reduction and Electrochemistry



Name Class Date In an oxidation-reduction reaction, What is the oxidation number assigned reduction is defined as the to manganese in KMnO<sub>4</sub>? A loss of protons sain of protons loss of electrons gain of electrons 3 A chemical cell is made up of two half-cells Which of the following aqueous solutions connected by an external conductor and a salt is the best conductor of electricity? bridge. The function of the salt bridge is to 5 **PREVIEW** D Please Sign In or Sign Up to download the printable version of this worksheet 7 supplied by the cell A Fe(s)  $\rightarrow$  Fe<sup>2+</sup> (aq) + 2e<sup>-</sup> A decreases **B**  $Fe^{2+}$  (aq)  $\to$   $Fe(s) + 2e^{-}$ **B** increase **C** Fe(s) +  $2e^- \rightarrow Fe^{2+}$  (aq) D Fe2 (aq) + 2e remains th same 9 h an oxidation-reduction reaction, the 14C decays to oxidation number of the oxidizing agent protons in the nucleus A decreases A decreases B **B** increases **B** increases c remains the same c remains the same