



Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

1 Which **reaction** occurs at the **anode** in **electrochemical cells** and in **electrolytic cells**?

- A reduction, only
- B oxidation, only
- C both reduction and oxidation
- D neither reduction nor oxidation

2 Which statement describes the **redox reaction** that occurs when an object is **electroplated**?

- A It is spontaneous and requires an electric current.
- B It is spontaneous and produces an electric current.
- C It is nonspontaneous and requires an electric current.
- D It is nonspontaneous and produces an electric current.

3 Which type of reaction produces **electrical energy** in a nickel-cadmium battery?

A redox

4 Which **metal** can be produced only by the **electrolysis of its fused salt**?

A Ag

5

**PREVIEW**

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As the  $\text{Cl}^-$  is oxidized, the **oxidation number** of chlorine will

- A decrease
- B increase
- C remain the same

- A anode, which is negatively charged
- B anode, which is positively charged
- C cathode, which is negatively charged
- D cathode, which is positively charged

9 Which **metals** are obtained by **electrolysis** of their **fused salts**?

- A K and Ca
- B K and Cr
- C Cu and Zn
- D Cu and Hg



10 Which **reaction** is an example of **oxidation-reduction**?

- A  $\text{KOH} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2\text{O}$
- B  $2\text{KCl} \rightarrow 2\text{K} + \text{Cl}_2$
- C  $\text{BaCl}_2 + \text{K}_2\text{SO}_4 \rightarrow 2\text{KCl} + \text{BaSO}_4$
- D  $\text{KCl} + \text{AgNO}_3 \rightarrow \text{AgCl} + \text{KNO}_3$



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