

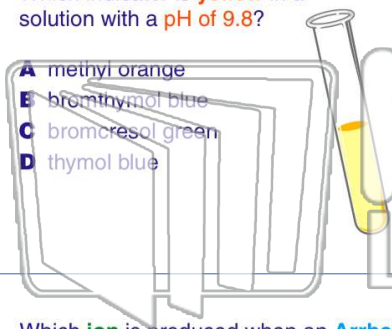


Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

1

Which indicator is **yellow** in a solution with a **pH of 9.8**?

- A methyl orange
- B bromthymol blue
- C bromocresol green
- D thymol blue



2

When 50. milliliters of an **HNO<sub>3</sub>** solution is **exactly neutralized** by 150 milliliters of a 0.50 M solution of **KOH**, what is the **concentration** of **HNO<sub>3</sub>**?

- A 1.0 M
- B 1.5 M
- C 3.0 M
- D 0.5 M

3

Which **ion** is produced when an **Arrhenius base** is dissolved in water?

A H<sup>+</sup> as the only positive ion in solution

4

If 5.0 milliliters of a 0.20 M HCl solution is required to **neutralize** exactly 10. milliliters of NaOH, what is the **concentration** of the NaOH?

5

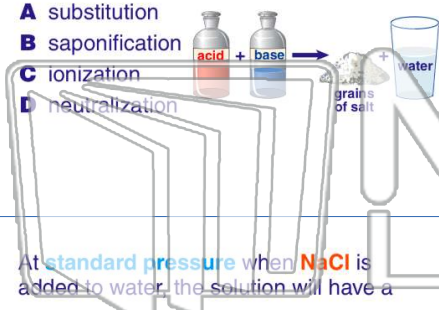


## PREVIEW

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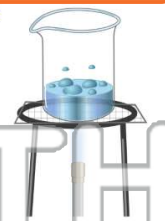
7

- A substitution
- B saponification
- C ionization
- D neutralization



between reacting particles

- A decreases
- B increases
- C remains the same



9

At **standard pressure** when **NaCl** is added to water, the solution will have a

- A higher freezing point and a lower boiling point than water
- B higher freezing point and a higher boiling point than water
- C lower freezing point and a higher boiling point than water
- D lower freezing point and a lower boiling point than water

10

Which **0.1 M solution** contains an **electrolyte**?

- A C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>(aq)
- B CH<sub>3</sub>COOH(aq)
- C CH<sub>3</sub>OH(aq)
- D CH<sub>3</sub>OCH<sub>3</sub>(aq)

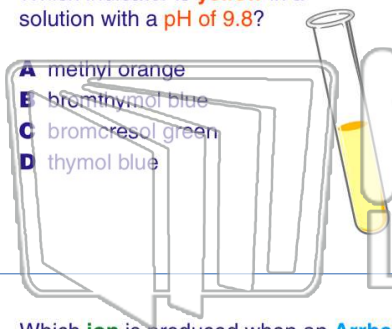


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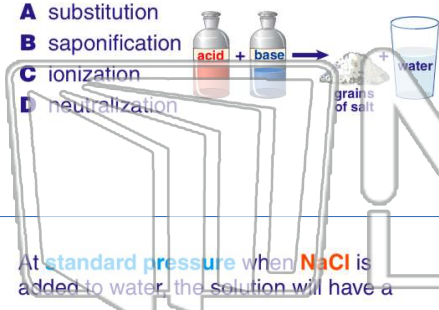
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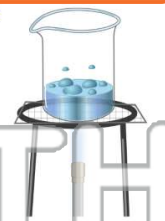
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