

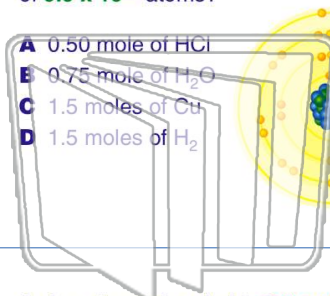


Name _____ Class _____ Date _____

1

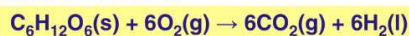
Which sample contains a total of 9.0×10^{23} atoms?

- A 0.50 mole of HCl
- B 0.75 mole of H_2O
- C 1.5 moles of Cu
- D 1.5 moles of H_2



2

Given the reaction:



How many moles of $C_6H_{12}O_6(s)$ are needed to produce 24 moles of carbon dioxide?

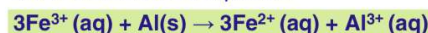
- A 1.0 mole
- B 12 moles
- C 24 moles
- D 4.0 moles

3

A closed container holds 3.0 moles of CO_2 gas at STP. What is the total number of moles of Ne(g) that can be placed in a container of the same size at STP?

4

Given the balanced equation:



What is the total number of moles of

5



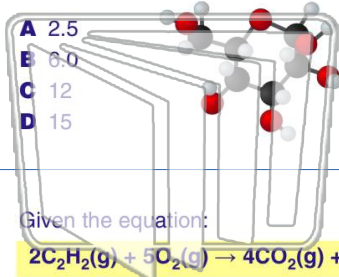
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7

What is the total number of moles of water needed to make 2.5 moles of $C_6H_{12}O_6$?

- A 2.5
- B 6.0
- C 12
- D 15



liter of solution?

- A 1.0 M
- B 2.0 M
- C 0.25 M
- D 0.50 M



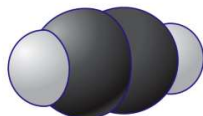
9

Given the equation:



How many moles of oxygen are required to react completely with 1.0 mole of C_2H_2 ?

- A 2.5
- B 2.0
- C 5.0
- D 10



10

What is the molarity of a solution of NaOH if 2 liters of the solution contains 4 moles of NaOH?

- A 0.5 M
- B 2 M
- C 8 M
- D 80 M

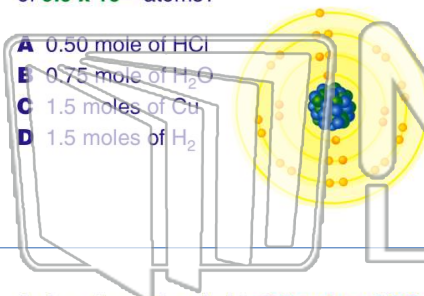




Name _____ Class _____ Date _____

- 1 Which sample contains a total of 9.0×10^{23} atoms?

A 0.50 mole of HCl
 B 0.75 mole of H₂O
 C 1.5 moles of Cu
 D 1.5 moles of H₂



- 2 Given the reaction:
 $C_6H_{12}O_6(s) + 6O_2(g) \rightarrow 6CO_2(g) + 6H_2(l)$

How many moles of $C_6H_{12}O_6(s)$ are needed to produce 24 moles of carbon dioxide?

A 1.0 mole C 24 moles
 B 12 moles D 4.0 moles

- 3 A closed container holds 3.0 moles of CO₂ gas at STP. What is the total number of moles of Ne(g) that can be placed in a container of the same volume at STP?

- 4 Given the balanced equation:
 $3Fe^{3+}(aq) + Al(s) \rightarrow 3Fe^{2+}(aq) + Al^{3+}(aq)$
 What is the total number of moles of

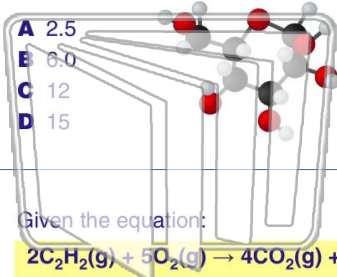


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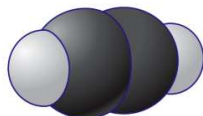
- 7 What is the total number of moles of water needed to make 2.5 moles of $C_6H_{12}O_6$?

A 2.5
 B 6.0
 C 12
 D 15



- 9 Given the equation:
 $2C_2H_2(g) + 5O_2(g) \rightarrow 4CO_2(g) + 2H_2O(g)$
 How many moles of oxygen are required to react completely with 1.0 mole of C_2H_2 ?

A 2.5
 B 2.0
 C 5.0
 D 10



liter of solution?

A 1.0 M
 B 2.0 M
 C 0.25 M
 D 0.50 M



- 10 What is the molarity of a solution of NaOH if 2 liters of the solution contains 4 moles of NaOH?

A 0.5 M
 B 2 M
 C 8 M
 D 80 M

