



Name _____ Class _____ Date _____

1 Compared to an atom of **potassium**, an atom of **calcium** has a

- A larger radius and lower reactivity
- B larger radius and higher reactivity
- C smaller radius and lower reactivity
- D smaller radius and higher reactivity

2 Elements in the Periodic Table are **arranged** according to their

- A atomic number
- B atomic mass
- C relative activity
- D relative size

3 Which **Group 15** element exists as a **diatomic molecule** at STP?

A phosphorus

7
N Nitrogen
15
P

4 Atoms of **metals** tend to

- A lose electrons and form negative ions

4
Be Beryllium
12
Mg Magnesium

5



PREVIEW

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7

- A colorless ions in solution, multiple positive oxidation states
- B colorless ions in solution, multiple negative oxidation states
- C colored ions in solution, multiple positive oxidation states
- D colored ions in solution, multiple negative oxidation states

- A decrease in size as it forms a positive ion
- B increase in size as it forms a positive ion
- C decrease in size as it forms a negative ion
- D increase in size as it forms a negative ion

9 The properties of **carbon** are expected to be **most similar** to those of

- A boron
- B aluminum
- C silicon
- D phosphorus

					18
5	6	7	8	9	10
S	C	N	O	F	Ne
Sulfur	Carbon	Nitrogen	Oxygen	Fluorine	Neon
13	14	15	16	17	18
Al	Si	P	S	Cl	Ar
Aluminum	Silicon	Phosphorus	Sulfur	Chlorine	Argon
31	32	33	34	35	36
Ga	Ge	As	Se	Br	Kr
Gallium	Germanium	Arsenic	Selenium	Bromine	Krypton
48	50	51	52	53	54
In	Sn	Sb	Te	I	Xe
Indium	Tin	Antimony	Tellurium	Iodine	Xenon
81	82	83	84	85	86
Tl	Pb	Bi	Po	At	Rn
Thallium	Lead	Bismuth	Polonium	Astatine	Radon

10 Which statement **best** describes **Group 2** elements as they are considered in order from **top to bottom** of the Periodic Table?

- A increases, and the number of valence electrons increases.
- B increases, and the number of valence electrons remains the same.
- C remains the same, and the number of valence electrons increases.
- D remains the same, and the number of valence electrons decreases.



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