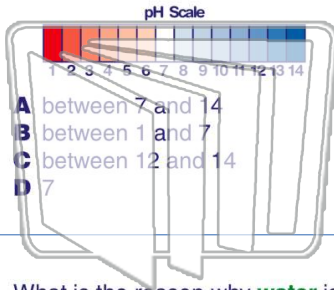




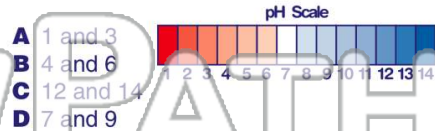
Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

1 Using the **pH scale** shown below, determine the pH of **acid rain**.



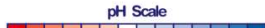
- A between 7 and 14
- B between 1 and 7
- C between 12 and 14
- D 7

2 **Weaker bases** usually have a pH between \_\_\_\_\_.



- A 1 and 3
- B 4 and 6
- C 12 and 14
- D 7 and 9

3 What is the reason why **water** is considered **neutral** on the pH scale?



4 A student wants to neutralize a **base** with a pH of **13**. To do this she should add an **acid** with a pH of \_\_\_\_\_.

5



## PREVIEW

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7



- A only water is produced
- B only salt is produced
- C the pH goes above 7
- D HCl and NaOH lose their chemical properties

9 In the reaction below, what is the approximate **pH** of substance #1?



- ①      ②      ③      ④
- A 1-6
  - B 7-10
  - C 10-12
  - D 12-14

10 It is possible for **acid rain** to irritate a person's skin.

True or false?

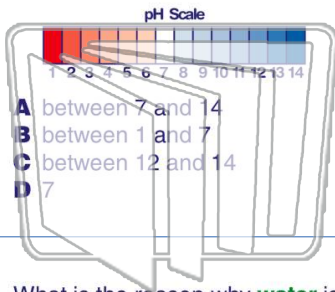
- A true
- B false



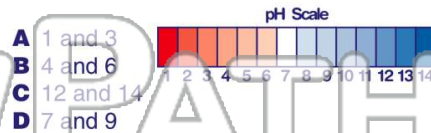


Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

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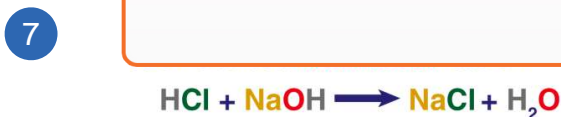


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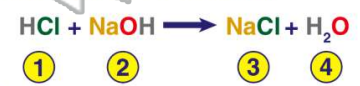
## PREVIEW

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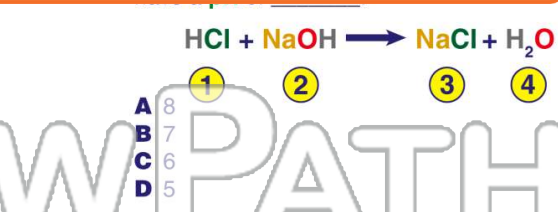


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It is possible for **acid rain** to irritate a person's skin. True or false?

- A true
- B false

