

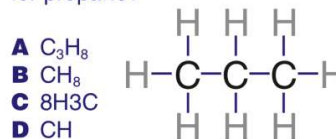


Name _____ Class _____ Date _____

1 How many atoms of **oxygen (O)** are in the **chemical formula** below?

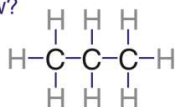


2 The diagram shown below is a **structural formula** for the gas **propane**. What would be the correct **chemical formula** for propane?



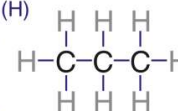
3 Carbon (C) and hydrogen (H) create **covalent bonds** to share **electrons** between them. What is the total number of electrons being shared between **carbon** and **hydrogen** in the molecule of propane shown below?

- A 8 C 16
B 11 D 32



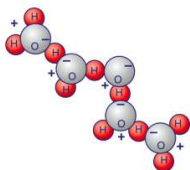
4 The **reason** why the bonds between carbon (C) and hydrogen (H) are **covalent** is _____.

- A C and H do not have electrons to give away
B C and H are both nonmetals
C C and H are both metalloids
D C and H can bond only by sharing electrons



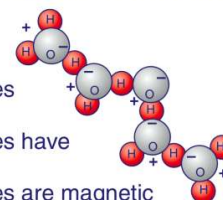
5 The diagram below represents the **structural formula** for several **water molecules**. How many **molecules** are represented?

- A 3
B 5
C 10
D 15



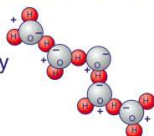
6 Explain the **arrangement** of the molecules in the diagram of water below.

- A it is a lucky arrangement
B water molecules are sticky
C water molecules have polarity
D water molecules are magnetic



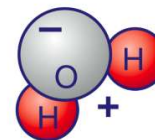
7 Water molecules have some **polarity** in their structure. Using the diagram below, notice how the ends of H atoms (**positive**) and O atoms (**negative**) meet. This property is the **reason** that _____.

- A water forms droplets easily
B water boils easily
C ice melts easily
D water has color



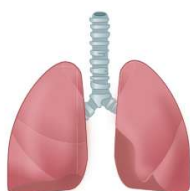
8 Since H atoms are slightly **positive** and the O atoms are slightly **negative**, molecules of water have **polarity**. This makes water a **good** _____.

- A solute
B refrigerant
C solvent
D magnet



9 The chemical formula for the **oxygen** that we breathe is _____.

- A O
B O_2
C O_4
D 2O



10 How many **molecules of oxygen** are represented in the chemical formula below?

- A 2
B 4
C 6
D 18





Name _____ Class _____ Date _____

1 How many atoms of **oxygen (O)** are in the **chemical formula** below?

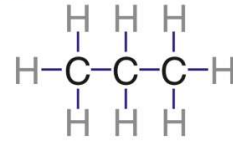
- A 4
- B 5
- C 7
- D 12



(D)

2 The diagram shown below is a **structural formula** for the gas **propane**. What would be the correct **chemical formula** for propane?

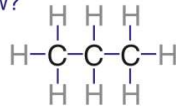
- A C_3H_8
- B CH_8
- C $8\text{H}3\text{C}$
- D CH



(A)

3 Carbon (C) and hydrogen (H) create **covalent bonds** to share **electrons** between them. What is the total number of electrons being shared between **carbon** and **hydrogen** in the molecule of propane shown below?

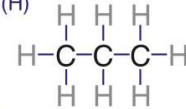
- A 8
- B 11
- C 16
- D 32



(C)

4 The **reason** why the bonds between carbon (C) and hydrogen (H) are **covalent** is _____.

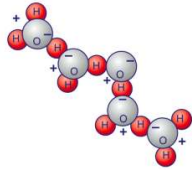
- A C and H do not have electrons to give away
- B C and H are both nonmetals
- C C and H are both metalloids
- D C and H can bond only by sharing electrons



(D)

5 The diagram below represents the **structural formula** for several **water molecules**. How many **molecules** are represented?

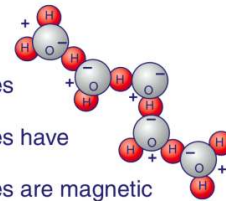
- A 3
- B 5
- C 10
- D 15



(B)

6 Explain the **arrangement** of the molecules in the diagram of water below.

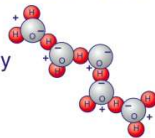
- A it is a lucky arrangement
- B water molecules are sticky
- C water molecules have polarity
- D water molecules are magnetic



(C)

7 Water molecules have some **polarity** in their structure. Using the diagram below, notice how the ends of H atoms (**positive**) and O atoms (**negative**) meet. This property is the **reason** that _____.

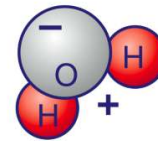
- A water forms droplets easily
- B water boils easily
- C ice melts easily
- D water has color



(A)

8 Since H atoms are slightly **positive** and the O atoms are slightly **negative**, molecules of water have **polarity**. This makes water a **good** _____.

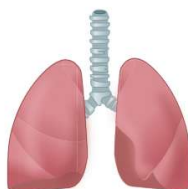
- A solute
- B refrigerant
- C solvent
- D magnet



(C)

9 The chemical formula for the **oxygen** that we breathe is _____.

- A O
- B O_2
- C O_4
- D 20



(B)

10 How many **molecules of oxygen** are represented in the chemical formula below?

- A 2
- B 4
- C 6
- D 18



(B)