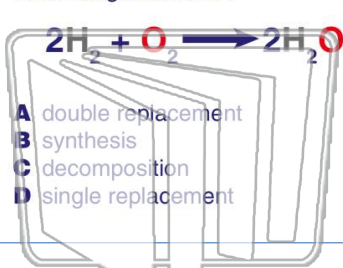


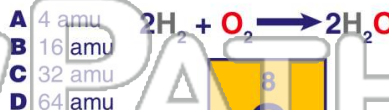


Name _____ Class _____ Date _____

1 What **type** of **chemical reaction** is shown in the diagram below?



2 The total mass of **oxygen** on **each side** of this chemical equation is _____.



3 In the reaction shown below, the **mass** of the reactant on the left **equals** the mass of the reactants on the right. What law is illustrated by this fact?

4 What **type** of **chemical reaction** is shown in the reaction below?



PREVIEW

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7 (moving quickly) the match?

- A** to provide a reactant
B to soften the match
C to provide activation energy
D to add another reactant

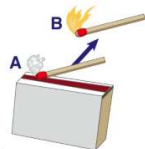


- A** light and heat
B sound
C change of shape
D change in mass



9 What **two terms** accurately describe the two steps shown here from **A** to **B**?

- A** endothermic, then exothermic
B only physical changes for both steps
C exothermic for both steps
D activation energy, then exothermic



10 The equation below shows how CaCO_3 decomposes into CaO and CO_2 . How many **grams** of CaCO_3 must be broken down in order to produce **112 grams** of CaO and **88 grams** of CO_2 ?





Name _____ Class _____ Date _____

1 What **type** of **chemical reaction** is shown in the diagram below?



- A double replacement
- B synthesis
- C decomposition
- D single replacement

2 The total mass of **oxygen** on **each side** of this chemical equation is _____.



- A 4 amu
- B 16 amu
- C 32 amu
- D 64 amu

8
O
16

3 In the reaction shown below, the **mass** of the reactant on the left **equals** the mass of the reactants on the right. What law is illustrated by this fact?

4 What **type** of **chemical reaction** is shown in the reaction below?



PREVIEW

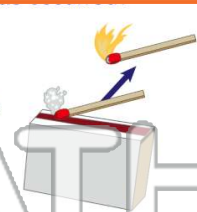
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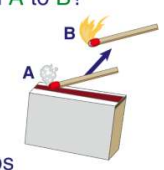


- A light and heat
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- A endothermic, then exothermic
- B only physical changes for both steps
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10 The equation below shows how CaCO_3 decomposes into CaO and CO_2 . How many **grams** of CaCO_3 must be broken down in order to produce **112 grams** of CaO and **88 grams** of CO_2 ?

- A 112
- B 24
- C 200
- D 224

$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$		
20	8	6
Ca	O	C
40	16	12