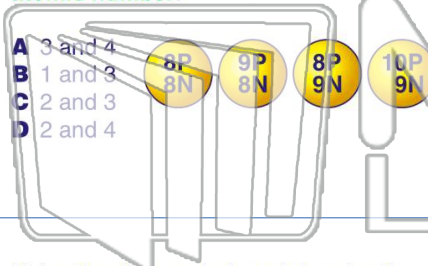




Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

1 Using the diagram below, determine which **two atoms** have the same **atomic number**.



2 Using the diagram below, determine which **two atoms** are **isotopes** of each other.



3 Using the diagram below, determine the **mass number** of atom #4.

4 Using the diagram below, determine which **area** of the **atom** is **positively** charged.

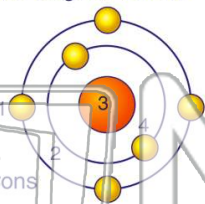


## PREVIEW

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7

A protons and electrons  
B only protons  
C neutrons and electrons  
D protons and neutrons



A the same atomic number and same atomic mass  
B different atomic numbers but the same atomic mass  
C the same atomic number but different atomic masses  
D the same number of neutrons

9 This diagram illustrates a **cloud** of **electrons** around the **nucleus**. This represents \_\_\_\_\_ model of the atom.

- A Bohr's
- B Rutherford's
- C Dalton's
- D the modern



10 The diagram below is a modern model of an atom. Why are **no electron orbits** drawn?

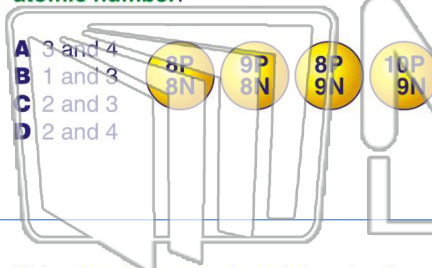
- A electrons move in many different directions
- B orbits are there but not visible
- C the orbits are not perfectly round
- D the electrons have been drawn into the nucleus





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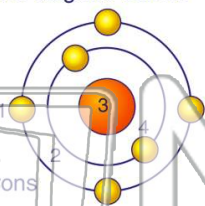


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