



Properties of Atoms

Name _____ Class _____ Date _____

1 Which two atoms have the **same atomic number**. Circle them.

8P 8N	9P 8N	8P 9N	10P 9N
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6 Two **atoms** of the **same element** must have the same _____.

- number of protons
- mass number
- number of neutrons

4 Be Beryllium	7 N Nitrogen
11 Na Sodium	8 O Oxygen
19 K Potassium	53 I Iodine

2 **Isotopes** are atoms that have _____. Circle the answer letter.

- the same atomic number and same atomic mass
- different atomic numbers but the same atomic mass

7 **Electron orbits** are sometimes referred to as **energy levels**. What is the explanation for this?

- Electrons have no charge.
- Electrons need energy to stay near _____

3 T
a
o
a
b
c



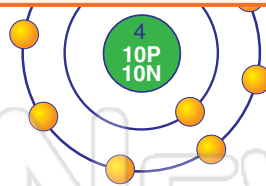
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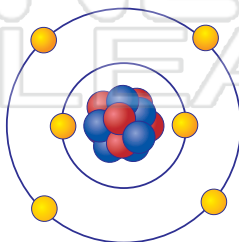
4 P
a
n

What is the **overall charge** of this atom?

- positive neutral
negative static

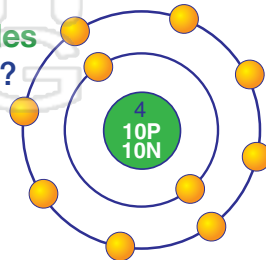


5 Fill in the answer.
An **atom** is the **smallest part** of a(n) _____.



10 How many **atomic particles** are shown in the diagram?

_____ protons _____ neutrons
_____ electrons





Properties of Atoms - Answer Key

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- a. number of protons
- b. mass number
- c. number of neutrons

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- b. different atomic numbers but the same atomic mass

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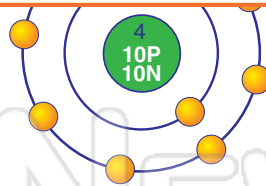


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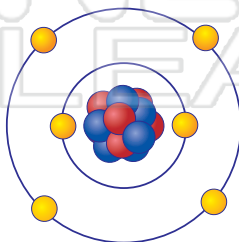
4 P
a
n
What is the **overall charge** of this atom?

- positive
- negative
- neutral
- static



- a. Bonded atoms of different elements make up compounds.
- b. Elements bond to make up atoms.
- c. Compounds are made from only one element.

5 Fill in the answer.
An **atom** is the **smallest part** of a(n) _____
element _____.



10 How many **atomic particles** are shown in the diagram?

10 protons 10 neutrons
10 electrons

